



Fact Sheet

Titratable Acidity

A way of measuring acidity



This work is licensed under the Creative Commons Attribution-NonCommercial-NoDerivatives 4.0 International License. To view a copy of this license, visit <http://creativecommons.org/licenses/by-nc-nd/4.0/> or send a letter to Creative Commons, PO Box 1866, Mountain View, CA 94042, USA



Equipment

- Erlenmeyer flask (conical flask) (100 ml)
- Burette (°SH/°D/°Th)
- Beaker
- Syringe or pipet (2 ml)
- Syringe or pipet (20 ml)

Reagents

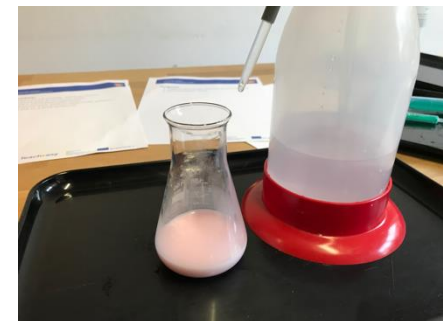
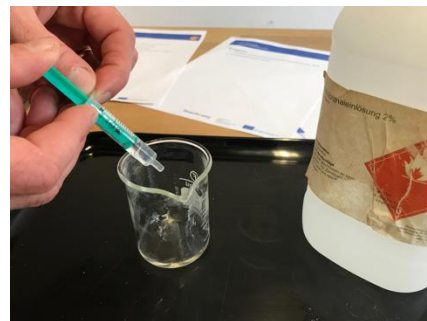
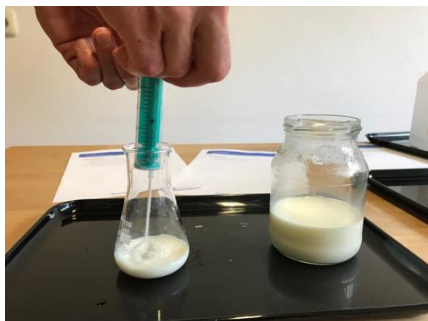
- Sodium hydroxide solution
- Phenolphthalein





Procedure

- Pipet milk, whey or starter culture into Erlenmeyer flask.
- Add Phenolphthalein.
- Fill buret with Sodium hydroxide solution.
- Titrate to a pink endpoint.
- Read the buret.





Difference in application

Procedure	Soxhlet-Henkel (° SH)	° Dornic (°D)	° Thörner (°Th)	Percentage of lactic acid (% l.a.)
Pipet milk into Erlenmeyer flask	25 ml milk	10 ml milk	10 ml milk + 20 ml distilled water	10 ml milk
Add Phenolphthalein	1 ml of Phenolphthalein (2%)	3-4 drops Phenolphthalein (1%)	5 drops of Phenolphthalein (5%)	3-4 drops of Phenolphthalein (1%)
Fill buret with Sodium hydroxide solution	1/4 N Sodium hydroxide solution	1/9 N Sodium hydroxide solution	1/10 N Sodium hydroxide solution	1/9 N Sodium hydroxide solution
Read the buret	consumption of base in ml = ° SH	consumption of base in ml = ° D	consumption of base in ml = ° Th	consumption of base in ml / 10 = % l.a.